

# MEASURE AND MANAGE

## Guide to Lowering the pH of Water

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The main factor influencing the alkalinity of water is the bicarbonate level. The more bicarbonate ions present, the more acid is required to lower the pH to a desired level.

Use the attached graph to determine the amount of acid required to lower the pH of 1000 liters of water to a desired value.

### HOW TO USE THE GRAPH?

1. Determine the bicarbonate level of your water.
2. From the graph choose the **titration** curve closest to the bicarbonate level of your water.
3. From the Y-axis choose the pH you wish to attain.
4. Read the corresponding amount of acid required from the X-axis.

### NOTES:

1. The amount of acid in mL on the X-axis is based on treating 100L of water.
2. Exercise care when adding the acid below pH 6.0, as it will drop dramatically with small increments of acid added.
3. Theoretically there are different titration curves for each bicarbonate level that lie between the ones presented. All these curves will have the same shape adjacent to each other.

### WATER TITRATION CURVES @ DIFFERENT HCO LEVELS (1000 LITRES)

